

Identifying Ions in a Salt:

Identify the ions that make up the salts. Ignore solubility. Put correct charges. Balance the "equation."

1. $\text{CuSO}_4 (aq) \rightarrow$
2. $\text{Fe}_2\text{O}_3 (aq) \rightarrow$
3. $\text{Mg}_3\text{N}_2 (aq) \rightarrow$
4. $\text{Al}_2\text{O}_3 (aq) \rightarrow$
5. $\text{Co}(\text{NO}_3)_2 (aq) \rightarrow$
6. $\text{Na}_3\text{PO}_4 (aq) \rightarrow$
7. $(\text{NH}_4)_2\text{SO}_4 (aq) \rightarrow$
8. $\text{NH}_4\text{Cl} (aq) \rightarrow$
9. $\text{Al}_2(\text{CO}_3)_3 (aq) \rightarrow$
10. $\text{Mg}(\text{OH})_2 (aq) \rightarrow$
11. $\text{KOH} (aq) \rightarrow$
12. $\text{H}_2\text{SO}_4 (aq) \rightarrow$
13. $\text{HCl} (aq) \rightarrow$

Determining Solubility of a Salt:

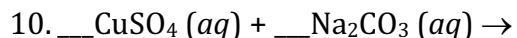
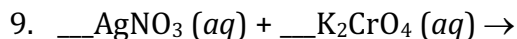
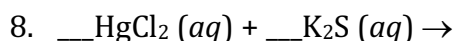
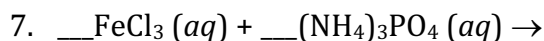
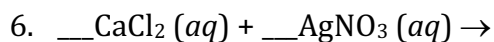
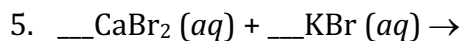
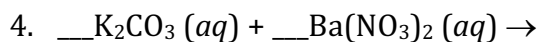
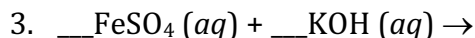
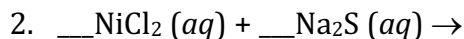
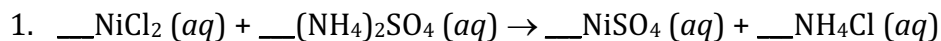
Use the solubility chart to determine if each salt is soluble or insoluble.

1. KNO_3 :
2. PbSO_4 :
3. KOH :
4. MgSO_4 :
5. FePO_4 :
6. Nickel (II) Hydroxide :
7. Sodium Chloride :
8. Barium Nitrate :
9. Ammonium Bromide :
10. Magnesium Hydroxide :

Mixing Salts: Molecular Equations

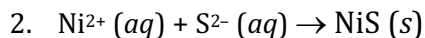
Determine the possible products and balance the equations.

Determine if each product is soluble (aq, aqueous) or an insoluble precipitate (s, solid).

**Mixing Salts: Net Ionic Equations**

For each of the molecular equations in the previous section, write the net ionic equation.

1. No net ionic equation



3.

4.

5.

6.

7.

8.

9.

10.